

Exam Points

Chemical Bonding ionic covalent & metallic:

- You must have understanding of the formulas of compound ions eg sulfate, hydroxide, nitrate, carbonate and ammonium.
- You must have understanding of covalent bond and its kinds like single, double and triple covalent bond.
- You should be able to represent a covalent bond using a line and co-ordinate bond using an arrow.
- You must have understanding of ions how these are formed and charges on the ions.
- You should be able to predict the charge on a simple ion using the position of the element in the Periodic Table.
- You should be able to construct formulas for ionic compounds.
- You must have understanding of the chemical bonds and intermolecular forces.
- You should be able to relate the melting point and conductivity of materials to the type of structure and the bonding present.
- You should be able to explain the energy changes associated with changes of state.
- You must have understanding of diagrams to represent the structures involving specified numbers of particles.
- You must have understanding of the four types of crystal structure.
 - ionic
 - metallic
 - macromolecular (giant covalent)
 - molecular.

I am Sorry !!!!!



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