CHEMISTRY ONLINE
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CHEMISTRY

MULTIPLE CHOICE - 2

CHEMICAL BONDING

Chemical Bonding – 2

1) Helping concepts

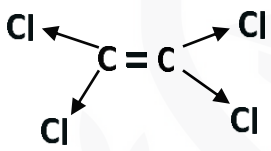
As number of carbon atoms of the homologues increase, the van der Waals' forces increases so that the homologues vaporize less readily. Hence, vapour pressure decrease.

2) Helping concepts

CCl4 exists as discrete molecules and its molecules are non – polar. The forces operating between its molecules are id – id interactions.

3) *Helping concept*

Although there is a permanent dipole in each C – Cl bonds, the effect of each cancels one another vectorically due to the symmetrical distribution of the 4 Cl atoms.



4) *Helping concept*

By definition, lattice energy is the energy released when 1 mole of an ionic compound is formed from its constituent gaseous ions (infinitely apart) combine together.

5) *Helping Concepts*

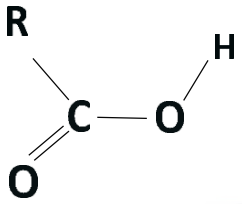
$$L.E. \propto -\left|\frac{q_+ q_-}{r_+ + r_-}\right|$$

CsCl has the largest r+ and r- among the ionic compounds. Hence, it has the least exothermic L.E.

6) *Helping Concept*

All the 5 compounds are isomers of hexane but (D) is unbranched. It has the greatest surface area of contact for VDW interaction and hence, the highest boiling point.

7) ***Helping Concept***



There are only 3 regions of electron clouds around c. To minimize electronic repulsion, they are directed in a trigonal planar manner and the bond angles are about 120^0.

8) ***Helping Concept***

In diamond and graphite, covalent bonds operate between carbon atoms and the extension of these bonds throughout lattices gives rise to macromolecular structure.

9) ***Helping Concept***

$$[NC \rightarrow Ag \leftarrow CN]^-$$

The 2 CN- ligands arrange themselves as far away from each other as possible to minimize electronic repulsion. Hence it is linear (bond angle = 180^0)

10) ***Helping Concept***

Dry HCl does not dissociate in non – polar solvents such as methylbenzene, Since there is no free ions or electrons, it does not conduct electricity.

I am Sorry !!!!!

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